

Chemistry Notes – Acids and Bases

- 1) Acids
 - a) Release _____ hydrogen ions (_____) in water
 - b) In water these cations combine to form _____ ions
 - c) Acids are _____, can conduct electricity, and can be corrosive.
- 2) Examples of Acids
 - a) _____ Acid
 - b) _____ Acid
 - c) _____ Acid
 - d) Nitric Acid
- 3) The strength of an acid is due to its ability to release the _____ from it.
- 4) Bases
 - a) Bases accept _____ to make water.
 - b) Bases are commonly associated with a _____.
 - c) Bases are used in soaps and are also corrosive like acids.
- 5) pH- Hydrogen Ion Potential
 - a) How acidic or basic a solution is
 - b) Ranges from _____
 - c) Neutral solutions are at a pH of _____
 - d) pH is directly related to the _____ ion and _____ ion concentrations in a solution.
 - i) pH is an _____ scale
 - ii) Something with a pH of 1 is _____ times more acidic than a solution at pH 3.
 - iii) A solution at pH 14 is 1×10^{14} more basic than a solution at pH _____.
- 6) pH indicators- indicate the pH of a solution
 - a) litmus paper
 - b) natural indicators- _____ juice
- 7) Neutralization- The interaction between _____ to form a _____ causing the pH to become closer to pH 7.